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Chapter 10 Supplemental Problems Balance the following chemical equations. 1)  $\text{SnS}_2(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{SnO}_2(\text{s}) + \text{SO}_2(\text{g})$  2)  $\text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$  3)  $\text{Al}(\text{s}) + \text{HCl}(\text{aq}) \rightarrow \text{AlCl}_3(\text{aq}) + \text{H}_2(\text{g})$  4)  $\text{CoCO}_3(\text{s}) \rightarrow \text{CoO}(\text{s}) + \text{CO}_2(\text{g})$  Write a balanced equation for each of the following reactions, substituting symbols and formulas for names.

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hydrogen-1 are in a 1.007-g sample? Recall that 1.007 amu is the mass of one atom of hydrogen-1. Round your answer to two significant digits. 1.007 g 1 H-1 atom \_\_ 1 ...

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