

Chapter 12 Basics Of Chemistry

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chemistry. • Describe oxidation and reduction (redox) reactions. • Discuss the different forms of matter: elements, compounds, and mixtures. • Explain the difference between solutions, suspensions, and emulsions. • Explain pH and the pH scale.

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organic chemistry. The study of substances that contain the element carbon. oxidation. A chemical reaction that combines a substance with oxygen to produce an oxide. oxidation-reduction. Also known as redox; a chemical reaction in which the oxidizing agent is reduced and the reducing agent is oxidized. oxidizing agent.

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A chemical reaction that combines a substance with oxygen to produce an oxide. oxidation-reduction. (aka redox) a chemical reaction in which the oxidizing agent is reduced (by losing oxygen) and the reducing agent is oxidized (by gaining oxygen). oxidizing agent. Substance that releases oxygen.

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A stable physical mixture of 2 or more substances. ALKANOLAMINES. Alkaline substances used to neutralize acids or raise the PH of many hair products. EMULSION. An unstable physical mixture of 2 or more immiscible substances, plus a special ingredient called a "emulsifier".

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Organic Chemistry provides a comprehensive discussion of the basic principles of organic chemistry in their relation to a host of other fields in both physical and biological sciences. This book is written based on the premise that there are no shortcuts in organic chemistry, and that understanding and mastery cannot be achieved without devoting adequate time and attention to the theories and concepts of the discipline. It lays emphasis on connecting the basic principles of organic chemistry to real world challenges that require analysis, not just recall. This text covers topics ranging from structure and bonding in organic compounds to functional groups and their properties; identification of functional groups by infrared spectroscopy; organic reaction mechanisms; structures and reactions of alkanes and cycloalkanes; nucleophilic substitution and elimination reactions; conjugated alkenes and allylic systems; electrophilic aromatic substitution; carboxylic acids; and synthetic polymers. Throughout the book, principles logically evolve from one to the next, from the simplest to the most complex examples, with abundant connections between the text and real world applications. There are extensive examples of biological relevance, along with a chapter on organometallic chemistry not found in other standard references. This book will be of interest to chemists, life scientists, food scientists, pharmacists, and students in the physical and life sciences. Contains extensive examples of biological relevance Includes an important chapter on organometallic chemistry not found in other standard references Extended, illustrated glossary Appendices on thermodynamics, kinetics, and transition state theory

Emphasises on contemporary applications and an intuitive problem-solving approach that helps students discover the exciting potential of chemical science. This book incorporates fresh applications from the three major areas of modern research: materials, environmental chemistry, and biological science.

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